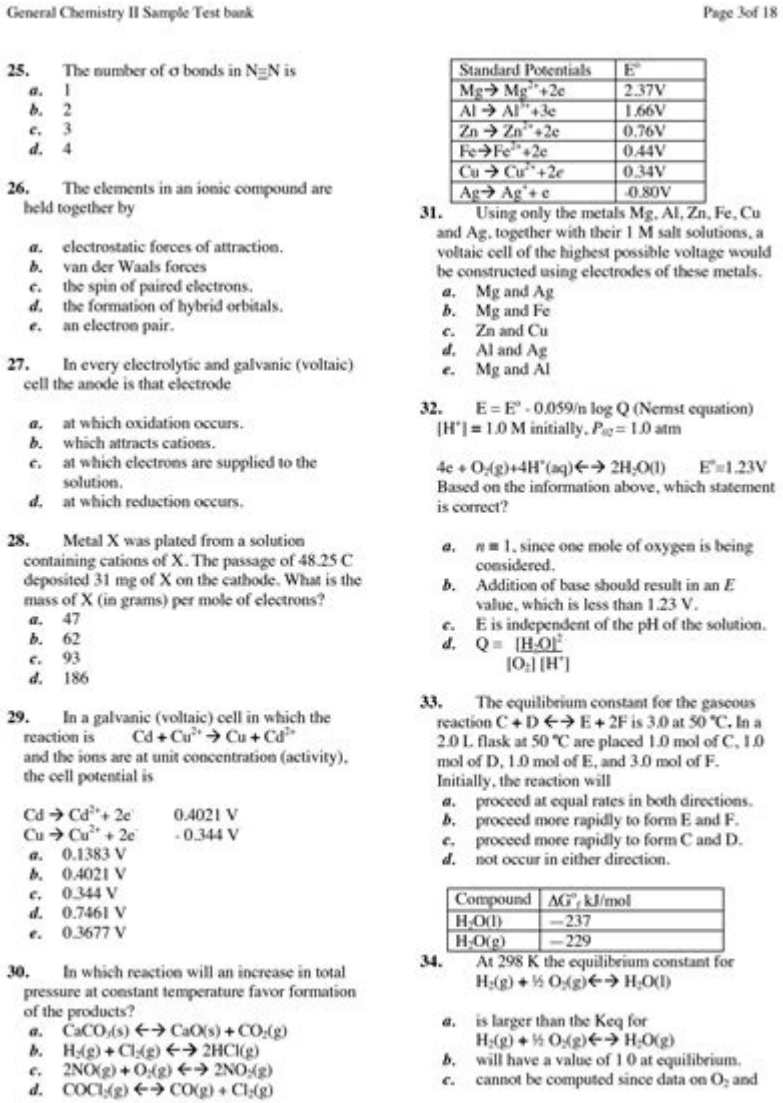
Acs General Chemistry 2 Exam



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25. The number of σ bonds in N≡N is
a. 1
b. 2
c. 3
d. 4

26. The elements in an ionic compound are held together by

a. electrostatic forces of attraction.
b. van der Waals forces
c. the spin of paired electrons.
d. the formation of hybrid orbitals.
e. an electron pair.

27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode

a. at which oxidation occurs.
b. which attracts cations.
c. at which electrons are supplied to the solution.
d. at which reduction occurs.

28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
a. 47
b. 62
c. 93
d. 186

29. In a galvanic (voltaic) cell in which the reaction is $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$ and the ions are at unit concentration (activity), the cell potential is

$Cd \rightarrow Cd^{2+} + 2e^-$ 0.4021 V
$Cu \rightarrow Cu^{2+} + 2e^-$ -0.344 V
a. 0.1383 V
b. 0.4021 V
c. 0.344 V
d. 0.7461 V
e. 0.3677 V

30. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
a. $CaCO_3(s) \leftrightarrow CaO(s) + CO_2(g)$
b. $H_2(g) + Cl_2(g) \leftrightarrow 2HCl(g)$
c. $2NO(g) + O_2(g) \leftrightarrow 2NO_2(g)$
d. $COCl_2(g) \leftrightarrow CO(g) + Cl_2(g)$

Standard Potentials	E°
$Mg \rightarrow Mg^{2+} + 2e$	2.37V
$Al \rightarrow Al^{3+} + 3e$	1.66V
$Zn \rightarrow Zn^{2+} + 2e$	0.76V
$Fe \rightarrow Fe^{2+} + 2e$	0.44V
$Cu \rightarrow Cu^{2+} + 2e$	0.34V
$Ag \rightarrow Ag^{+} + e$	-0.80V

31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
a. Mg and Ag
b. Mg and Fe
c. Zn and Cu
d. Al and Ag
e. Mg and Al

32. $E = E^\circ - 0.059/n \log Q$ (Nernst equation)
$[H^+] = 1.0$ M initially, $P_{O2} = 1.0$ atm

$4e + O_2(g) + 4H^+(aq) \leftrightarrow 2H_2O(l)$ $E^\circ = 1.23V$
Based on the information above, which statement is correct?

a. $n = 1$, since one mole of oxygen is being considered.
b. Addition of base should result in an E value, which is less than 1.23 V.
c. E is independent of the pH of the solution.
d. $Q = \frac{[H_2O]^2}{[O_2][H^+]}$

33. The equilibrium constant for the gaseous reaction $C + D \leftrightarrow E + 2F$ is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
a. proceed at equal rates in both directions.
b. proceed more rapidly to form E and F.
c. proceed more rapidly to form C and D.
d. not occur in either direction.

Compound	ΔG°_f, kJ/mol
$H_2O(l)$	−237
$H_2O(g)$	−229

34. At 298 K the equilibrium constant for $H_2(g) + \frac{1}{2} O_2(g) \leftrightarrow H_2O(l)$

a. is larger than the Keq for $H_2(g) + \frac{1}{2} O_2(g) \leftrightarrow H_2O(g)$
b. will have a value of 1 0 at equilibrium.
c. cannot be computed since data on O_2 and

Conquering the ACS General Chemistry 2 Exam: A Comprehensive Guide

Are you staring down the barrel of the ACS General Chemistry 2 exam, feeling overwhelmed and unsure of where to begin? You're not alone. This notoriously challenging exam requires not just a strong understanding of the material, but also effective study strategies and a clear exam-taking approach. This comprehensive guide will equip you with the tools and knowledge to confidently face the ACS General Chemistry 2 exam and achieve your desired results. We'll dissect the exam format, explore effective study techniques, and offer practical tips to maximize your performance.

Understanding the ACS General Chemistry 2 Exam Landscape

The ACS General Chemistry 2 exam is a standardized test designed to assess your comprehension of general chemistry principles. Unlike many other exams, it emphasizes problem-solving and critical thinking, demanding a deep understanding beyond simple memorization. Understanding the exam's structure is the first step to success.

Exam Format and Content:

The ACS General Chemistry 2 exam typically consists of multiple-choice questions covering a wide range of topics, including:

Thermodynamics: Enthalpy, entropy, Gibbs free energy, spontaneity, equilibrium constants.
Chemical Kinetics: Reaction rates, rate laws, activation energy, reaction mechanisms.
Chemical Equilibrium: Equilibrium constants, Le Chatelier's principle, solubility product.
Acid-Base Chemistry: pH, pKa, buffers, titrations.
Electrochemistry: Oxidation-reduction reactions, electrochemical cells, Nernst equation.
Nuclear Chemistry: Radioactivity, nuclear reactions, half-life.
Spectroscopy: Basic principles of various spectroscopic techniques (UV-Vis, IR, NMR).

The weighting of each topic can vary slightly from year to year, so consulting the most recent ACS exam specifications is crucial.

Time Management Strategies:

Effective time management is critical during the exam. Practice taking timed practice exams to develop a sense of your pacing and identify areas where you tend to spend too much time. Prioritize questions you find easier and return to the more challenging ones if time permits.

Mastering the Material: Effective Study Techniques for ACS General Chemistry 2

Simply reading the textbook isn't enough to conquer this exam. A multi-faceted approach is essential.

Active Recall and Practice Problems:

Active recall, the process of retrieving information from memory without looking at your notes, is one of the most effective study techniques. Regularly test yourself using flashcards, practice problems, and past exams. The ACS website and various textbooks offer ample practice problems.

Focus on Conceptual Understanding:

While memorizing formulas is important, a deep understanding of the underlying concepts is crucial. Don't just memorize equations; understand why they work and how they apply to different scenarios. Work through example problems step-by-step, ensuring you understand each stage of the calculation.

Seek Clarification:

Don't hesitate to ask for help if you're struggling with a particular concept. Utilize office hours with your professor, form study groups with classmates, or seek assistance from a tutor.

Utilizing Online Resources:

Numerous online resources can supplement your studies. Explore reputable websites offering practice problems, videos explaining challenging concepts, and interactive simulations.

Exam Day Strategies: Tips for Success

Preparation is only half the battle. On exam day, ensure you're well-rested, have eaten a nutritious breakfast, and have all necessary materials.

Read Carefully:

Pay close attention to the wording of each question. Often, slight changes in phrasing can significantly alter the answer.

Eliminate Incorrect Answers:

If you're unsure of the correct answer, try eliminating obviously incorrect choices to increase your chances of selecting the correct one.

Review Your Work:

If time permits, review your answers before submitting the exam. Look for careless errors and ensure your calculations are accurate.

Conclusion: Achieving Success on the ACS General Chemistry 2 Exam

The ACS General Chemistry 2 exam is challenging, but with dedicated effort, the right study strategies, and a confident approach, you can achieve your desired score. Remember to focus on conceptual understanding, practice consistently, and manage your time effectively. Good luck!

Frequently Asked Questions (FAQs)

Q1: What is the passing score for the ACS General Chemistry 2 exam?

A1: There isn't a publicly stated "passing score." The score is typically reported as a percentile ranking, comparing your performance to other students who have taken the exam. Your institution will determine what score constitutes a passing grade.

Q2: Are calculators allowed during the ACS General Chemistry 2 exam?

A2: Yes, typically scientific calculators are permitted, but programmable calculators and those with communication capabilities are usually prohibited. Check your exam's specific regulations.

Q3: How many questions are on the ACS General Chemistry 2 exam?

A3: The exact number of questions can vary slightly, but it generally falls within the range of 70-80 multiple-choice questions.

Q4: What type of reference materials are allowed?

A4: Generally, no reference materials are permitted during the ACS General Chemistry 2 exam. The exam is designed to test your understanding of fundamental principles.

Q5: When should I start studying for the ACS General Chemistry 2 exam?

A5: The earlier, the better. A thorough review of the material typically requires several weeks or even months, especially if you haven't encountered some of the topics recently. Starting early allows you to pace your studying effectively and address any gaps in your understanding.

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technology issues, with hands-on experimentation and data collection. It contains 30 experiments to aid the understanding of the scientific method and the role that science plays in addressing societal issues. Experiments use microscale equipment (wellplates and Beral-type pipets) and common materials. Project-type and cooperative/collaborative laboratory experiments are included.

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